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Experiment 1

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I-1 CHEM 2115 Experiment # 1

INTRODUCTION TO THE LABORATORY

OBJECTIVES: The goals of this session include (1) an introduction to the

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general chemistry laboratory and CHEM 2115 course policies and procedures, (2) safety instruction, (3) checking into lab lockers, (4) learning how to use

CHEM 2115 Experiment # 1 - Stockton University

Chem 2115 is a co-requisite of Chem 2110. Students will perform experiments

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in measurement, stoichiometry, solutions, chemical equilibrium, and acids & bases; demonstrating principles learned in Chem 2110.

CHEM 2115: General Chemistry I Laboratory - Spring 2014

Download class data for this experiment from Blackboard, and calculate the value

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of R each data set. After you have calculated all of the R values, inspect the data. If any points are obviously far off from the others do not include them in the average.

CHEM 2115 - Stockton University

Chem 2115 is a co-requisite of Chem 2110. Students will perform experiments

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in measurement, stoichiometry, solutions, chemical equilibrium, and acids & bases; demonstrating principles learned in Chem 2110. Course Policies
The following policies have been adopted by the Chemistry Program and apply to all sections of Chem I Laboratory. 1.

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CHEM 2115: General Chemistry I Laboratory - Spring 2018

Chemistry Program, School of Natural
Sciences and Mathematics 101 Vera
King Farris Drive, Galloway, NJ 08205
CHEM 2115: Chem I Lab Experiment 9 —
Additional Instructions Task Overview:
You will use measured experimental
data collected by previous students to

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determine the value of the ideal gas constant, R . For the report you will: 1.

CHEM 2115: Chem I Lab Experiment 9 – Additional Instructions

CHEM 2115 Experiment 1 Introduction to the Analytical Balance OBJECTIVES: To acquire skills needed for the quick and accurate determination of the mass of a

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sample using an analytical balance.

INTRODUCTION: The ability to determine mass accurately is fundamental not only to analytical chemistry but also to all of modern science.

CHEM 2115 Experiment 1 Introduction to the Analytical Balance

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CHEM 2115 CHEMISTRY I LAB . School:
Stockton College of New Jersey
(Stockton University) * Professor: Carisa
Lautner, {[professorsList]} Carisa
Lautner ... Experiment 1-Intro-
ReportV2.docx. 3 pages.
8-Iron_Report-1.docx Stockton University
CHEMISTRY I LAB CHEM 2115 - Spring
2014 ...

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CHEM 2115 : CHEMISTRY I LAB - Stockton University

1 CHEM 2115 LAB REPORT Experiment
#1 Mass, Volume and Density Chem I
Lab Name Jordanna Mastroddi Section #
Station # A4 Date 01/24/17 1. Mass of
object next to the balance: Observed
mass 3.414g Posted mass 3.415g 2.

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Mass of 10 mL graduated cylinder: 3.
Volume in graduated cylinders: 4.

chem lab report exp 1 - 1CHEM 2115 LAB REPORT Experiment#1 ...

CHEM 2115 Experiment Two Gravimetric
Determination of Sulfate in Seawater
Objective The concentration of sulfate
ion in seawater will be determined

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gravimetrically by precipitation with barium chloride. Text Reference: Rubinson and Rubinson, Contemporary Chemical Analysis, Chapter 10.
Introduction

CHEM 2115 Experiment Two
Gravimetric Determination of ...
CHEM 2115 Experiment Three

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Determination Of The Composition Of A Multi-Component Mixture By Spectrophotometric Analysis Objective: The composition of a three-component mixture will be assayed by dissolution of the soluble components in water and measurement of their individual concentrations by visible spectrophotometry. Text Reference:

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CHEM 2115 Experiment Three Determination Of The ...

CHEM 2115 Experiment Six Separation of a Mixture by Ion-Exchange OBJECTIVE The objective of this experiment is to determine the composition of a mixture of metal salts. The mixture components are separated with ion chromatography

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and quantified by ... 1:1 complexes with EDTA ($\log K_{\text{ZnEDTA}} = 16.50$, $\log K_{\text{NiEDTA}} = 18.62$). PROCEDURE: 1. Accurately ...

CHEM 2115 Experiment Six
Separation of a Mixture by Ion ...
CHEM 2115 Experiment Five Ionization
of Phosphoric Acid OBJECTIVE The

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objective of this experiment is to learn the proper use and calibration of a pH meter and to use this instrument in determining the ionization constants for phosphoric acid. ... 1. As part of last week's experiment, you prepared a liter of ~0.2 N sodium hydroxide ...

CHEM 2115 Experiment Five

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Ionization of Phosphoric Acid ...

CHEM 2115 LAB REPORT Experiment #1

Mass, Volume and Density Chem I Lab

Name Section # Station # Date

Observed mass Written in mg Written in

kg Observed mass on milligram balance

Observed mass on centigram balance 1.

Mass of object next to the balance:

Observed mass Posted mass 2.

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1-Intro-Report-1 (1).docx - CHEM 2115 LAB REPORT ...

View 9-Gas-Law-Report.docx from CHEM 2115 at Stockton University. 1CHEM 2115 Lab Report Experiment #9 Ideal Gas Law Chem I Lab on # 006 Station # A2 Date 11/29/2016 I. Data Trial 1 Trial 2 Trial

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9-Gas-Law-Report.docx - 1CHEM

2115 Lab Report Experiment#9 ...

Question: CHEM 2115 Limiting Reagent

Lab Report #6 Chem I Lab Name

Mikhayla D'Anne Section# Station# A5

Date 1/4/19 Overall Equation: $\text{CaCl}_2(\text{aq})$

+ $\text{K}_2\text{C}_2\text{O}_4(\text{aq}) = \text{CaC}_2\text{O}_4(\text{s}) + 2$

$\text{KCl}(\text{aq})$ Net Ionic Equation: $\text{Ca}^{2+}(\text{aq}) +$

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CO. (aq) + H₂O (l) = CaCO₃ (s) + H₂O (l)
Data & Calculations : CaCl₂ K₂C₂O₄+H₂O
Concentration Of Solution 0.75 Mol/L
Volume Of Solution 15 ml Moles ...

Solved: CHEM 2115 Limiting Reagent Lab Report ... - Chegg.com

View Lab Report - 1 filled maybe from
CHEM 2115 at Stockton University.

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CHEM 2115 LAB REPORT Experiment #1
Name Mass, Volume and Density Section
1. Mass of object next to the balance:
Observed

**1 filled maybe - CHEM 2115 LAB
REPORT Experiment#1 Name ...**

CHEM 2115 Experiment Seven
Extraction of a Copper Chelate

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OBJECTIVE The objectives of this experiment are: (1) to measure the concentration of copper complex in both aqueous and nonaqueous samples using a spectrophotometer. (2) to evaluate the efficiency of the extraction of copper from aqueous solutions into

CHEM 2115 Experiment Seven

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Extraction of a Copper Chelate ...

Name Chem 2115 Experiment Date
Section GAS BEHAVIOR Trial 1 I. Data
Mass of test tube and contents 58.726 g
59.601 g before heating 58.619 g
58.726 g afterheating Masa loss 231.8
ml 259.6 mL 25.3 deg C 25.6 deg C
29.79 in Hg Volume of Co, collected, V
Temperature of Co,,T Atmospheric

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pressure Water vapor pressure at Co,
temperature 29.80 in Hg IL Calculations
Mass of Co, collected Moles of Co ...

Solved: Name Chem 2115 Experiment Date Section GAS BEHAVIO ...

The ideal gas law expresses $PV=nRT$
where in this experiment P stands for

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the pressure of gas, V is volume, n is the number of gas moles, T is the temperature given in Kelvins, and R the ideal gas constant given as $8.3144598 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ or $0.08206 \text{ L}\cdot\text{atm}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ (depending on our units being used).

**Lab Report 9 - CHEM 1100 General
Chemistry I - CSULA - StuDocu**

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Experimental Data Collection: The data collection procedure for this experiment involves two main steps: 1) construction of a calibration curve by measuring the absorbances of five standard solutions of iron, and 2) measurement of the absorbance of the analyte iron solution and determination of the corresponding concentration (using the calibration

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curve).

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